



A highly oxygen reduction reaction active and CO₂ durable high-entropy cathode for solid oxide fuel cells

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ABSTRACT

One big obstacle for the oxygen reduction reaction (ORR) electrode in solid oxide fuel cells (SOFCs) is the poor reaction activity and fast degradations caused by CO₂ poisoning. Here we report our design of an active A-site Ca-rich high-entropy Pr_{0.1875}Ba_{0.1875}Sr_{0.1875}La_{0.1875}Ca_{0.25}CoO_{3-δ} (PBSLC₂₅C) electrode, guided by the O *p*-band theory. When applied as a cathode in solid oxide fuel cells (SOFCs), it demonstrates high ORR activity and excellent CO₂ tolerance under realistic operating conditions. Ni-YSZ-based anode-supported cells with PBSLC₂₅C cathodes demonstrate excellent peak power densities of 1.14 W cm⁻², 1.04 W cm⁻², and 0.77 W cm⁻² in the air with 1%, 5%, and 10% CO₂, respectively, at 750 °C. The engineered high-entropy PBSLC₂₅C effectively diminishes the CO₂ poisoning effect and maintains active surfaces for fast oxygen exchange, as confirmed by the cell durability test in air containing CO₂ (5 and 10 vol%), Raman spectroscopy, and density functional theory calculations.

1. Introduction

Solid oxide fuel cells (SOFCs) offer promising prospects for the most efficient energy utilization and conversion using flexible fuels [1–3]. Rapid attention has been focused on the rational design of cathodes, which could contribute the most to the degradation of cell performance during the operation [4,5]. Oxygen reduction reaction (ORR) activity as an important parameter is of high merit in revealing the properties of functional cathodes [6–8]. However, the kinetics of ORR are complicated especially under realistic operating conditions with contaminations (e.g., CO₂, H₂O, or/and Cr), limited by one or more steps [9,10]. The electrochemical performance and stability of SOFCs are highly restricted by a largely increasing polarization resistance (*R_p*) of the cathode due to the dramatic poisoning from contaminations on the ORR activity [11,12]. Thus, future market penetration of the SOFC technique requires further innovation of active and durable cathodes to prolong the cell/system lifetime.

A descriptor (e.g., O *p*-band center, bulk vacancy energy, and charge-transfer energy) can demonstrate the key energetics underlying the ORR

activity of the cathodes for high-performance SOFCs [13–15]. Among these, the O *p*-band center as a theoretical descriptor can effectively realize the reliable prediction of ORR activity without the details of the complex electrode surface [13,16]. Adjusting the position of the O *p*-band center approaching the Fermi level can optimize the kinetics of the surface oxygen exchange and chemisorption energies. This method was extensively utilized for material designs of highly active catalysts in the field of electro-catalysis [17,18]. Even under elevated temperatures, the perovskite oxides with the disordered structure can be computationally investigated by employing special quasirandom structures (SQS) [19]. A descriptor of O *p*-band center is utilized effectively as a practical theoretical tool for developing optimal cathode catalysts in SOFCs.

The screened electrodes on the basis of the oxygen *p*-band center can enhance the long-term durability and intrinsic tolerance against contaminants poisoning [20,21]. According to the theoretical calculations of CO₂ adsorption energy, PrBa_{0.8}Ca_{0.2}CoO_{5+δ} (PBCC) perovskite materials have been currently investigated as a robust cathode for enhancing structural stability against CO₂ poisoning [11]. The PBCC electrode

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exhibited excellent durability in air containing CO₂, comprehensively verified by the stabilities of typical symmetrical cells (~1000 h) and single cells (~400 h) in CO₂-containing atmospheres [11]. A-site Ca substitution is chosen as a reasonable optimization strategy for the contaminant-tolerance electrodes with the guidance of the density functional theory (DFT) calculations [22]. SrCoO_{3-δ}-type perovskite materials have been extensively verified as high-performance cathodes for SOFCs due to their high O p-band values [23–25]. Theoretically, Sr²⁺ in the perovskite lattice can accelerate the intrinsic exchange kinetics of oxygen reduction and evolution reactions at the electrode surface [26, 27]. The activity of such SrCoO_{3-δ}-type electrodes is likely limited by the poor structural durability due to the severe surface Sr enrichment at elevated temperatures [28]. To maximize the role of Sr²⁺ cation in the SrCoO_{3-δ}-type perovskite materials, it is vital to improve the electrode durability and even tolerance against contaminants. The non-redox element (e.g., Nb, Ta, and Zr) doping strategy has been reported as an exceptional optimization of the structural stability of the electrode for a long-term operation [29]. In addition, the concept of entropy stabilization has been extensively studied for the active and robust cathodes via effectively restraining the Sr²⁺ segregation [12,30,31]. The contribution of multivariant A-site cations (e.g., Sr, Ca, and Pr, etc.) can perform interesting synergistic effects in the system of the high-entropy perovskite oxides [30,32]

Inspired by those studies, herein we report a novel high-entropy perovskite oxide Pr_{0.1875}Ba_{0.1875}Sr_{0.1875}La_{0.1875}Ca_{0.25}CoO_{3-δ} (PBSLC₂₅C) with A-site Ca cation rich, displaying a boosted electrochemical performance and an excellent durability against CO₂ (up to 10%). The anode-supported cells with the PBSLC₂₅C cathode achieved remarkable peak power densities of 1.14 W cm⁻², 1.04 W cm⁻², and 0.77 W cm⁻² at 750 °C in the air with different concentrations (varies from 1%, 5%, and 10% CO₂) of CO₂, respectively. It was found that the main contribution to the increasing resistance during the CO₂ poisoning process could be the electrode reaction of the surface gas-related diffusion, as analyzed by the distribution of relaxation time (DRT). On the other hand, the polarization resistance of the PrBa_{0.8}Sr_{0.2}Co₂O_{6-δ} (PBSC) cathode increased dramatically from 0.021 (in the air) to 0.186 Ω cm² (in the air with 5% CO₂), due likely to the aggravated formation of SrCO₃ deposited on the electrode surface under the condition of a high CO₂ concentration. The durability of symmetrical cells (300 h) and single cells (100 h) in air containing 5% and 10% CO₂ strongly indicate that the high-entropy PBSLC₂₅C cathode exhibits robust surface characteristics for an enhanced CO₂-tolerance, as probed by Raman spectroscopy and DFT calculations.

2. Experimental section

2.1. Materials synthesis

The high-entropy perovskite of Pr_{0.1875}Ba_{0.1875}Sr_{0.1875}La_{0.1875}Ca_{0.25}CoO_{3-δ} (PBSLC₂₅C) was synthesized by a sol-gel complexing method. Based on the nominal composition of PBSLC₂₅C, Pr(NO₃)₃·6 H₂O, Ba(NO₃)₂, Sr(NO₃)₂, Ca(NO₃)₂, La(NO₃)₃, and Co(NO₃)₂·6 H₂O were stoichiometrically added into the deionized water with complexing agents of EDTA and citric acid. Ammonia solution was then added to adjust the pH of the solution. The solution was continuously stirred and heated at 90 °C to form a gel. Subsequently, the achieved gel was dried at 250 °C for 5 h in an oven. Finally, the precursors were pre-calcined at 950 °C for 2 h, and then sintered at 1000 °C for 5 h to obtain PBSLC₂₅C powders. Some samples of PrBa_{0.8}Sr_{0.2}Co₂O_{6-δ} (PBSC), PrBa_{0.8}Ca_{0.2}Co₂O_{6-δ} (PBCC), and Pr_{0.2}Ba_{0.2}Sr_{0.2}La_{0.2}Ca_{0.2}CoO_{3-δ} (PBSLCC) were prepared by the same method, and all calcined at 1000 °C for 5 h. The anode and electrolyte powders of NiO, Zr_{0.84}Y_{0.16}O_{2-δ} (YSZ), and Gd_{0.1}Ce_{0.9}O_{2-δ} (GDC) were purchased from H2-BANK. La_{0.8}Sr_{0.2}Ga_{0.8}Mg_{0.2}O_{3-δ} (LSGM) was synthesized by a solid-state reaction method. LaO, SrCO₃, MgO, and Ga₂O₃ as raw materials were mixed by ball milling for 12 h, and then calcined at 1250 °C for 5 h to obtain

final powders.

2.2. Cell preparation

The LSGM powder was dry-pressed into green pellets and then sintered at 1450 °C for 5 h to obtain the dense pellets. Dense LSGM pellets with a thickness of 400 μm were prepared for symmetrical tests. The anode-supported half cells were fabricated by a co-tape casting method, which has been reported in our previous studies [33,34]. To eliminate the adverse reactions between the cathode and YSZ electrolyte, the GDC buffer layer was drop-coated onto the NiO-YSZ|YSZ half cells, and then co-sintered at 1300 °C for 2 h. The solution of the GDC buffer layer is composed of GDC powder, ethyl cellulose, terpinol, and acetone with a weight ratio of 1:0.15:1.85:10. The cathodes were then screen-printed on the GDC buffer layer and then sintered at 1000 °C for 2 h. The as-fabricated symmetrical cells were also sintered at 1000 °C for 2 h for further tests. The cathode ink for the screen-printing was composed of the cathode powder, ethyl cellulose, and terpinol with a weight ratio of 1:0.04:0.76. The effective area of single cells was 0.2826 cm². The silver wires were used as current collectors.

2.3. Characterization and electrochemical tests

X-ray diffraction (XRD) measurements were conducted by a Bruker D8 Advance diffractometer (Bruker, Germany) using Cu Kα radiation (λ=0.15406 nm). The microstructure and morphology of samples were observed by scanning electron microscopy (SEM) (Hitachi SU8010, Japan) and transmission electron microscopy (TEM) (TF20). The electrical conductivity relaxation (ECR) technique was used for the D* and k* of electrode materials, and the green cathode powders were dry-pressed and then sintered at 1175 °C for 10 h to obtain the dense bar samples. An X-ray photoelectron spectrometer (XPS) (Thermo Scientific K-Alpha) was employed to identify elements in samples. Inductively coupled plasma (ICP) analysis (Agilent 5110) was conducted to elucidate the equimolarity of a high-entropy perovskite cathode. The iodine titration method was used to examine the oxygen vacancy concentrations of the samples, and the details of iodine titration have been reported [35,36]. In this work, 0.8 g PBSLC₂₅C powders were dry-pressed in a bar model (with a length of 20 mm and a width of 6 mm) at a pressure of 1 MPa, and then sintered at 1150 °C for 10 h in the air to obtain the dense bars for thermal expansion coefficient (TEC) measurements. The thermal expansion coefficient (TEC) of the sample was tested by the equipment of Netzsch DIL 402 in the air from room temperature to 1000 °C at the heating rate of 20 °C min⁻¹. The CO₂-temperature programmed desorption (TPD) was conducted by a mass spectrometer (HPR-20EGA), and the samples were all pre-treated in the 10% CO₂/air atmosphere. Raman spectroscopy (RenishawRM1000) was conducted using a wavelength of 514 nm laser to detect the carbonate or/and deposited carbon on the electrode surface. The electrochemical impedance of the electrodes was measured in the air, the air with different pCO₂, or different pO₂ by a potentiostat (PARSTAT MC 200), respectively. For tests of SOFC performance, the cells were sealed onto the alumina tube with ceramic paste (AREMCO 552), and then put inside the furnace and heated to the needed temperatures at the heating rate of 3 °C min⁻¹. The fuel electrode was flowed with humidified hydrogen (3% H₂O) at the rate of 60 ml min⁻¹, while the air electrode was exposed to air and air containing 1%, 5%, and 10% CO₂ at the total rate of 100 ml min⁻¹. The air and CO₂ gases were accurately controlled by two flow meters (KMG3–21) at concentrations of 1%, 5%, or 10%, respectively. The frequency range for all impedance measurements was between 100 mHz and 10 kHz with an AC amplitude of 10 mV. The I-V-P curves and EIS of single cells were tested using a potentiostat (PARSTAT MC 200) at 700–800 °C, and electrochemical impedance spectroscopic (EIS) measurements of symmetrical cells were also performed using a potentiostat (PARSTAT MC 200) at 650–850 °C.

2.4. Computational methods

Periodic density functional theory (DFT) calculations were carried out using the Vienna ab initio simulation package (VASP) [37,38] to examine the enhanced ORR activity and the CO₂ tolerance of high-entropy PBSLC₂₅C. The projector augmented wave (PAW) [39] method and the Perdew-Burke-Ernzerhof (PBE) [40] exchange-correlation functional was used with a kinetic energy cutoff of 415 eV. In particular, to accurately describe the strongly correlated electronic states of the Co cation at the B site [41], the Hubbard U correction with the U_{eff} value of 4.0 eV was applied [42]. The detailed computational methods are available in **Supplementary Note 1**.

3. Results and discussion

3.1. Structural analysis of the Ca-rich high-entropy PBSLC₂₅C

Fig. 1a schematically depicts the structure of the high-entropy PBSLC₂₅C. The actual atomic ratio of as-synthesized PBSLC₂₅C is close to the nominal stoichiometric ratio, as validated by the inductively coupled plasma mass spectrometry (ICP-MS) (Table S1). Shown in Fig. 1b is the room-temperature (RT) X-ray diffraction (XRD) pattern of as-synthesized PBSLC₂₅C powders sintered at 1000 °C for 5 h. The refinement demonstrated that Ca-rich PBSLC₂₅C has a cubic perovskite structure with the space group of $Pm\bar{3}m$ and lattice parameters of $a=3.8351(8)$ Å (a fitted factor of 1.16), similar to the original PBSLCC (space group of $Pm\bar{3}m$, $a=3.8500$ Å) (Table S2) [31]. The phase structural variations on the PBSLC₂₅C were evaluated by high-temperature XRD (HT-XRD) from RT to 800 °C (Fig. S1). The distinctive perovskite peaks at 31–35° shifted towards the low angles without the appearance

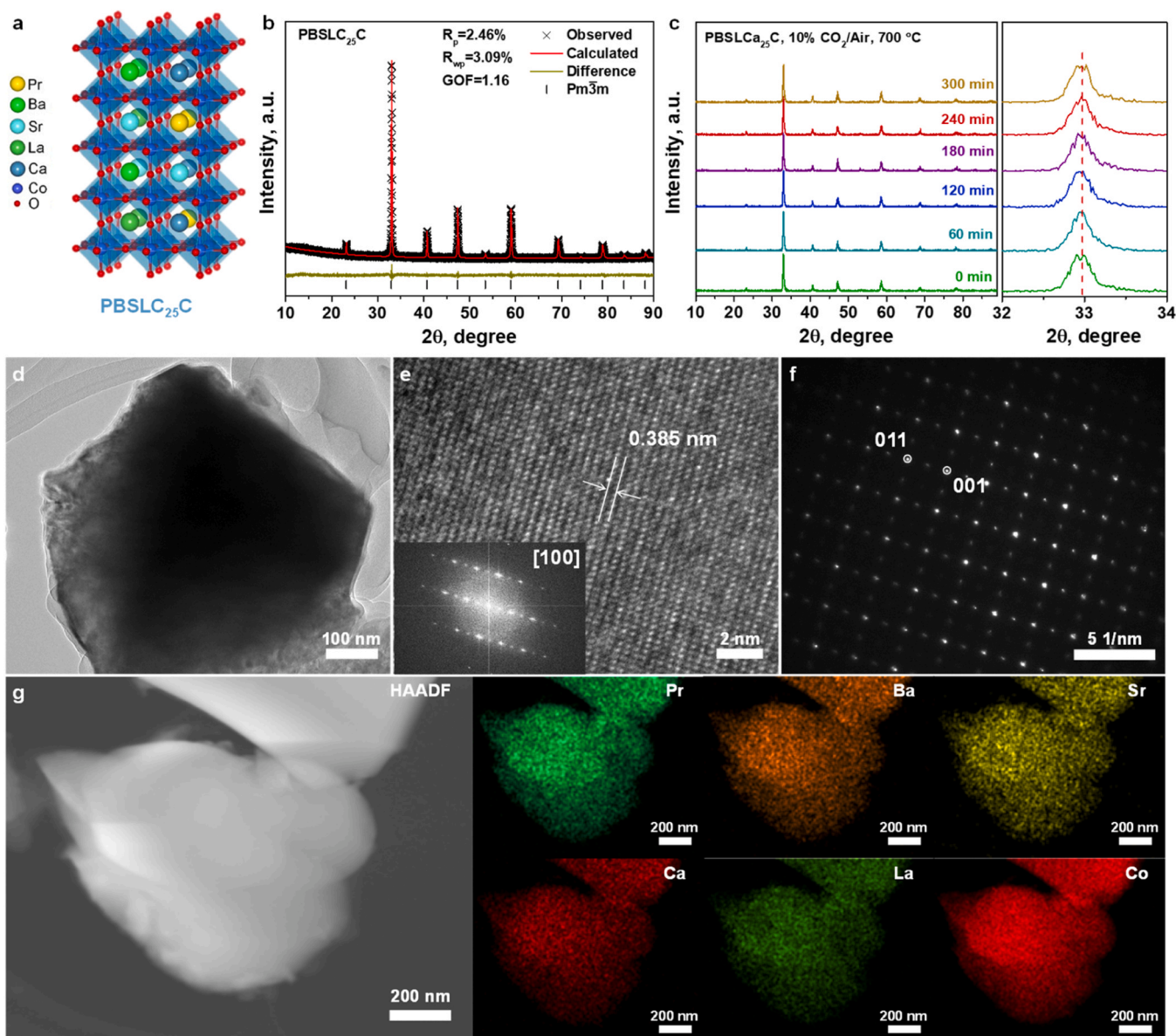


Fig. 1. Crystalline structure of PBSLC₂₅C powders. (a) Schematic of the non-equimolar PBSLC₂₅C. (b) Refined XRD pattern of as-synthesized PBSLC₂₅C powders. (c) In-situ XRD patterns of PBSLC₂₅C powders in the air containing 10% CO₂ at 700 °C and different processing times. (d) STEM and (e) high-resolution TEM images of a PBSLC₂₅C grain. The inset is the fast Fourier transform (FFT) pattern in (e). (f) A SAED image of the PBSLC₂₅C grain. (g) High-angle annular dark-field STEM image combined with the EDS mapping of the PBSLC₂₅C grain.

of secondary phases at elevated temperatures to 800 °C. To further examine the structural stability of PBSLC₂₅C, an in-situ XRD was performed in air containing 10% CO₂ at 700 °C with intervals of 60 min (Fig. 1c). No peaks representing carbonates and apparent shifts were observed during each poisoning processing time for a total of 300 min, confirming the excellent structural tolerance of PBSLC₂₅C against CO₂ contaminant. Shown in Fig. 1d is the scanning transmission microscopy (STEM) image of the PBSLC₂₅C grain. The space distance between the adjacent bright planes is about 3.85 Å, consistent with the (100) plane of PBSLC₂₅C from the XRD results (Fig. 1e). The inset (Fig. 1e) is the fast Fourier transform (FFT) based on the high-resolution TEM image, and the marked diffraction spots correspond to the planes of [011] and [001], as shown in the selected area electron diffraction (SAED) image of the PBSLC₂₅C (Fig. 1f). The element distributions of PBSLC₂₅C were also examined by the high-angle annular dark-field (HAADF) STEM and X-ray energy dispersive spectrum (EDS) (Fig. 1g). The six elements of Pr, Sr, Ba, Ca, La, and Co are homogeneously distributed in the selected high-entropy particle without Ca cation enrichments.

3.2. Electrochemical performance and theoretical interpretation of the highly active PBSLC₂₅C electrode

Shown in Fig. S2 are the electrochemical impedance spectra (EIS) of the PBSLC₂₅C cathode determined by symmetrical cells with a configuration of PBSLC₂₅C|LSGM|PBSLC₂₅C in ambient air at 650–850 °C. The LSGM electrolyte is a typical oxygen ion-conducting electrolyte with excellent chemical and thermal compatibility [51], which was applied to evaluate the electrochemical performance of electrodes using symmetrical cells in this work. Herein, LSGM was used for the symmetrical cells to avoid the use of a protective GDC barrier layer. The thermal expansion coefficient (TEC) of the PBSLC₂₅C above 500 °C is $25.1 \times 10^{-6} \text{ K}^{-1}$, and that below 500 °C is

$18.1 \times 10^{-6} \text{ K}^{-1}$. The relative high TEC above 500 °C may be ascribed to the thermal reduction of cobalt ions and thermally activated transition in the spin state of the Co *d*-orbital electrons [4]. The relatively high TEC could be reduced by mixing the cathodes with electrolyte phases [52] or negatively thermal expansion phases [4]. The interfacial R_p value ($0.037 \Omega \text{ cm}^2$ at 750 °C in the air) of PBSLC₂₅C is slightly lower than that ($0.043 \Omega \text{ cm}^2$ at 750 °C in the air) of PBSLCC at the same condition, which is also comparable to those of the high-performance cathodes with good stabilities reported recently (Fig. 2a and Table S3), including PrBa_{0.5}Sr_{0.5}Co_{1.5}Fe_{0.5}O_{5+δ} (PBSCF), [43] SrFe_{0.25}Ti_{0.25}Co_{0.25}Mn_{0.25}O_{3-δ} (SFTCM) [44], PrBaCo_{2/3}Fe_{2/3}Cu_{2/3}O_{5+δ} (PBCFC), [45] LaNi_{0.6}Co_{0.4}O_{3-δ} (LNCO) [46], Ba_{0.9}Co_{0.7}Fe_{0.2}Nb_{0.1}O_{3-δ} (BCFN), [47] Ba_{0.8}Gd_{0.8}Pr_{0.4}Co₂O_{6-δ} (BGPC), [48] Sm_{0.5}Sr_{0.5}CoO_{3-δ} (SSC)-GDC [49], SmBaCo₂O_{5+δ} (SBCO)-SDC [50], (La_{0.6}Sr_{0.4})_{0.95}Co_{0.2}Fe_{0.8}O_{3-δ} (LSCF) [48], Pr_{0.2}Ba_{0.2}Sr_{0.2}La_{0.2}Ca_{0.2}CoO_{3-δ} (PBSLCC) [31], and PBCC [31] composites. The higher electrocatalytic activity of the PBSLC₂₅C cathode may result from more oxygen vacancy concentrations at room temperature, as obtained by the typical iodine titration measurement (Fig. S4). The surface kinetics of the PBSLC₂₅C cathode were further investigated via electrochemical impedances tested under different oxygen partial pressures (p_{O_2}), and then the impedances were analyzed by the distribution of relaxation time (DRT) (Fig. S5 and Fig. 2b). According to the resistance contributions from each frequency, the electrode reaction at LF of total charge transfer (i.e., $\text{O}_{\text{ads}} + 2\text{e}^- + \text{V}_{\text{O}}^{\bullet\bullet} \rightarrow \text{O}_{\text{O}}^{\times}$) was highly associated with the key rate-limiting electrochemical step, as judged by the *n* value (reaction order) from the equation of $R_p = k(p_{\text{O}_2})^{-n}$ (Fig. 2c, Supplementary Note 2) [44,53,54]. A relatively stable R_p of $0.038 \Omega \text{ cm}^2$ was achieved within the stability test of 200 h in air at 750 °C (Figs. 2d and S6). The enhanced activity is attributed to the faster kinetics of surface oxygen exchange (k^*) and a higher coefficient of bulk oxygen transfer (D^*), compared to those of the PBSLCC reported recently (Fig. S7, Table S3).

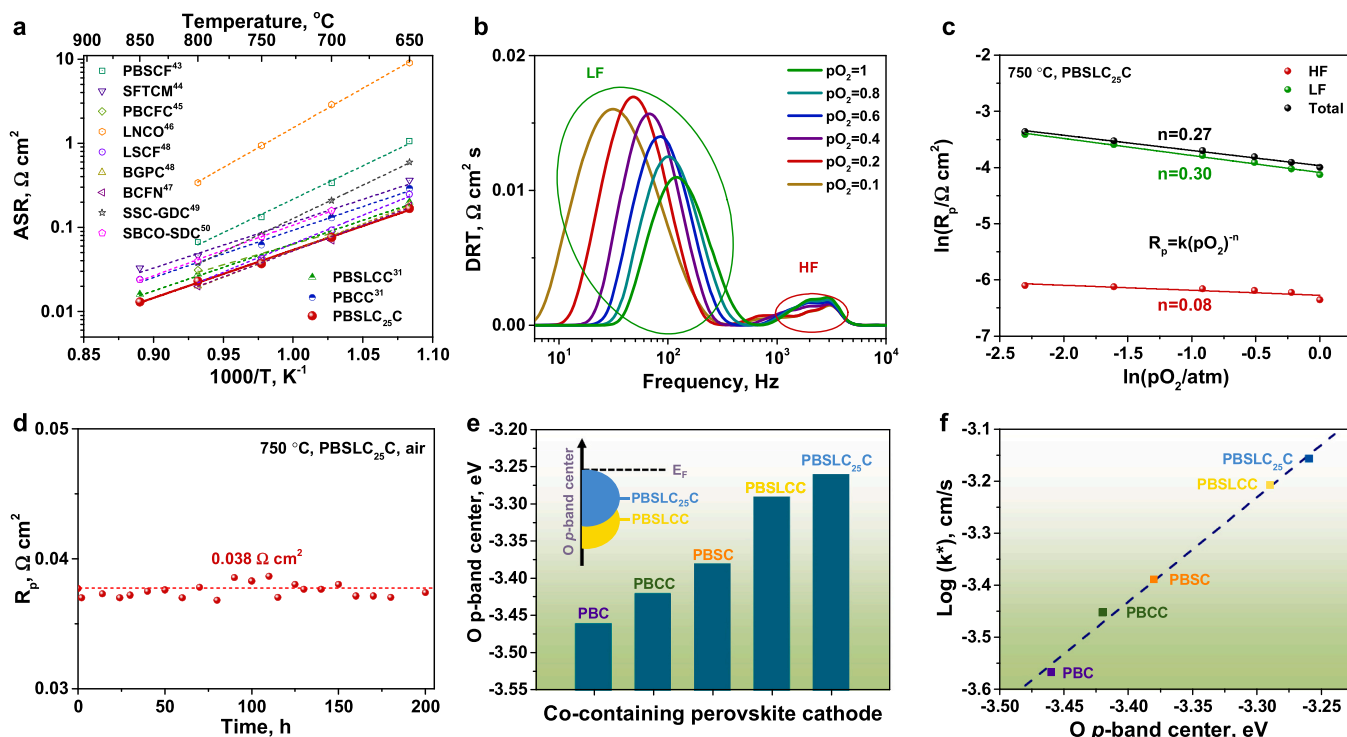


Fig. 2. Electrochemical performance and theoretical interpretation of the highly active PBSLC₂₅C electrode. (a) Arrhenius plots of ASRs of symmetrical cells with the PBSLC₂₅C electrode in comparison with many other reported electrode materials, including PBSCF,^[43] SFTCM,^[44] PBCFC,^[45] LNCO,^[46] BCFN,^[47] BGPC,^[48] SSC-GDC,^[49] SBCO-SDC,^[50] LSCF,^[48] PBSLCC,^[31] and PBCC^[31] composites; (b) DRT analysis of the PBSLC₂₅C electrode measured at different oxygen partial pressures (p_{O_2}) at 750 °C; (c) Dependence of polarization resistance (R_p) of PBSLC₂₅C at different frequency ranges and 750 °C as functions of p_{O_2} ; (d) Time dependence of R_p values of the PBSLC₂₅C electrode at 750 °C in the air; (e) Comparison of the bulk O *p*-band center of PBC, PBCC, PBSC, PBSLCC, and PBSLC₂₅C calculated using the hybrid HSE06 functional with the structures optimized at PBE + U; (f) Linear correlation of the experimental surface exchange coefficients (k^*) measured at 700 °C as a function of the bulk O *p*-band center for PBC, PBCC, PBSC, PBSLCC, and PBSLC₂₅C.

As shown in Fig. 2e, we further carried out density functional theory (DFT) calculations to elucidate the enhanced ORR activity and CO₂ tolerance of the high-entropy perovskite of PBSLC₂₅C. The bulk structures [51] used in this study are summarized in Table S4 and Fig. S8. As schematically depicted in Fig. 1a, additional Ca doping was applied in the equimolar high-entropy PBSLCC structure [51] by adding one more Ca ion to support the experimental observation, resulting in non-equimolar PBSLC₂₅C. Fig. S9 illustrates the strong hybridization of cations in the Co ions with the oxygen of PBSLCC and PBSLC₂₅C using the hybrid HSE06 functional.[55] Then, the bulk O *p*-band center, widely used as a descriptor for the rational design of SOFC cathode materials [13,56,57], was calculated as compiled in Table S5. It was found that more Ca doping in PBSLCC slightly shifts the bulk O *p*-band center to the Fermi level (−3.29 eV versus −3.26 eV). Moreover, we also examined that of PBSC as it was used for the comparative study of CO₂ tolerance by more Ca doping in the perovskite. Fig. 2e exhibits the apparent correlation of the perovskite-structured cathode materials, theoretically supporting the enhanced ORR activity [13,56,57] of the high-entropy PBSLC₂₅C cathode compared to PBC, PBCC, PBSC, and PBSLCC. More specifically, to evaluate the A-site entropy tuning of PBSLCC by additional Ca doping, we calculated the adsorption energy of oxygen by placing superoxide-like species [58] in the most stable, reactive CoO-terminated (001) plane (Table S5 and Fig. S10). Its trend aligns well with the O *p*-band behavior (Fig. S10f). In addition, as shown in Fig. 2f, we also confirmed that the first-principles-based O *p*-band centers are linearly correlated with their surface exchange coefficients (*k*^{*}) measured in this study (Fig. S11) and available in the literature [51]. We assumed that the additional Ca doping in equimolar, high-entropy PBSLCC may weaken oxygen adsorption to reach its maximum ORR activity by increasing the kinetics of diffusion and desorption processes.[59]

3.3. Robustness of the high-entropy PBSLC₂₅C electrode tested in air with CO₂

Apart from high electrocatalytic performance, the electrode durability against contaminants (such as CO₂) has attracted increasing concerns towards SOFC commercialization. When exposed to 1% CO₂/99% air, the *R*_p of the PBSLC₂₅C electrode reached up to 0.039 Ω cm² at 750 °C (Fig. S12), slightly higher than that (0.032 Ω cm² at 750 °C) of the PBSLC₂₅C electrode in the air. The electrode surface kinetics of the PBSLC₂₅C electrode in the air containing CO₂ were investigated by the DRT analysis on the electrochemical impedances (Figs. 3a and S13). It is found that the degradation on the kinetics of the electrode surface reaction did occur due likely to the CO₂ poisoning for ORR on the electrode surface. Based on the associated equation of $R_p = k(p_{CO_2})^m$, *m* = 0.12 for the peak at HF (Fig. S14) may indicate the transfer of O^{2−} through the LSGM electrolyte [60,61]. The gas-associated (CO₂ and/or O₂) diffusion processes of adsorption and dissociation (*m*_{LF} = 0.29, dependence on the variations of *p*_{CO₂}) [62,63] may be urgently in need of optimization when exposed to the air with CO₂ (Fig. S15, Supplementary Note 3). The excellent regenerative *R*_p stability of PBSLC₂₅C was verified by the well-recovered impedances, such as 0.044 Ω cm² in air/1% CO₂ (the first test for 140 h), 0.057 Ω cm² in air/5% CO₂ (the second test for 130 h), and then 0.048 Ω cm² in air/1% CO₂ (the third test for 80 h) (Fig. 3b, Supplementary Note 4). Finally, PBSLC₂₅C still showed the remarkable *R*_p stability of ~0.066 Ω cm² in air containing 10% CO₂ for the fourth stability test of 110 h (total 460 h). Comparatively, the *R*_p value of the PBSC electrode dramatically increased after the test with only 50 h exposure to air/5% CO₂ and air/10% CO₂. The *R*_p of the PBSC electrode is not recoverable to the pristine values (a two-fold increase) even in the air, much worse than the regenerative performance of the Ca-rich PBSLC₂₅C electrode (Fig. S15). The SEM images of the PBSLC₂₅C and PBSC electrodes after the 50-h exposure in air/5% CO₂ at 750 °C were

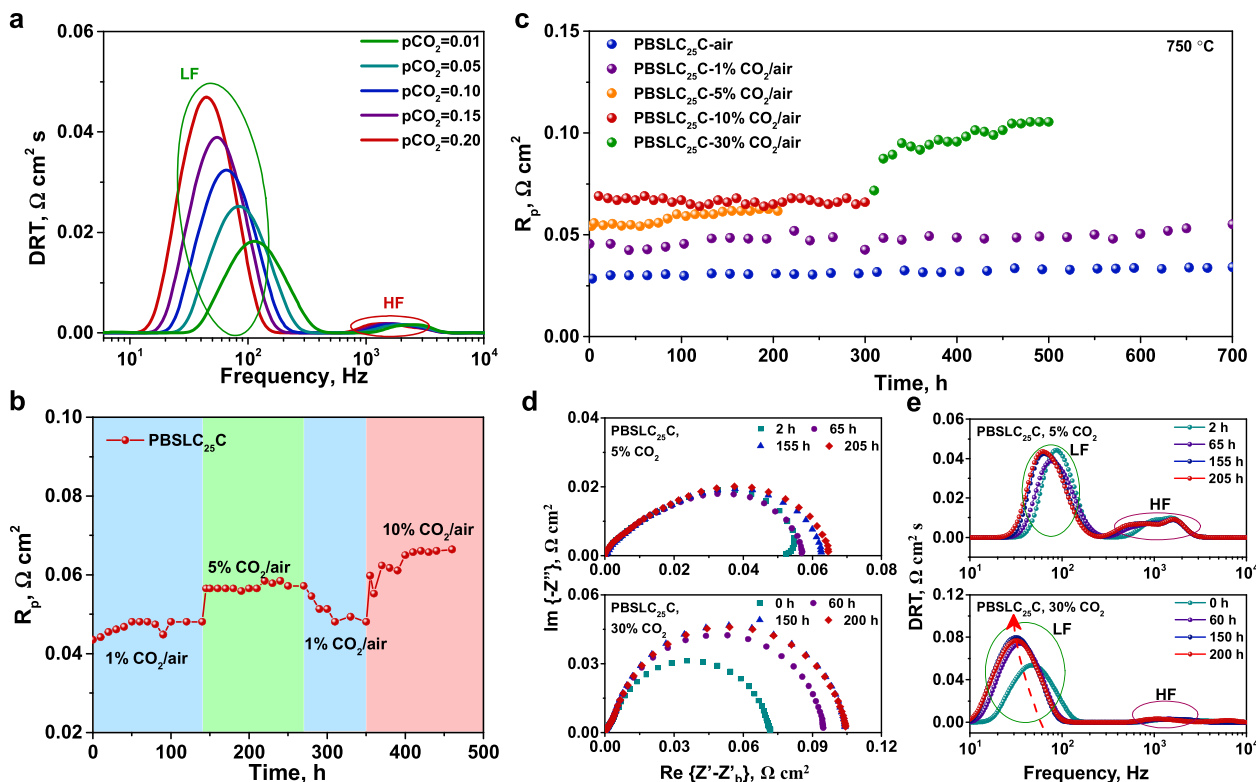


Fig. 3. Durability of PBSLC₂₅C in the air with various CO₂ concentrations. (a) DRT analysis of the PBSLC₂₅C electrode measured at different *p*CO₂ at 750 °C; (b) Regenerative *R*_p stability of PBSLC₂₅C when being switched among various CO₂ concentrations of 1%, 5%, and 10% at 750 °C; (c) Long-term durability of the PBSLC₂₅C cathode in air with different concentrations of CO₂ tested at 750 °C under OCV conditions; (d) EIS and (e) the corresponding DRT of the PBSLC₂₅C electrode tested at different times under different CO₂ concentrations: 5% (top) and 30% (bottom).

shown in Fig. S16. It is shown that the surface of the PBSC electrode was covered by abundant carbonate nanoparticles, rougher than that of the PBSLC₂₅C electrode and differing from the original PBSC.

Shown in Fig. 3c is the summary of the durability of the PBSLC₂₅C electrode in air with various concentrations of CO₂. The R_p of the PBSLC₂₅C electrode exceptionally remained stable at 0.034 and 0.045 $\Omega \text{ cm}^2$ when exposed to the air and the air containing 1% CO₂, respectively. Even when the CO₂ concentration increased to 5%, 10%, and 30%, the R_p of the PBSLC₂₅C electrode increased at a relatively low rate. As comparisons, the R_p of high-entropy PBSLCC and Sr-free PBCC electrodes increased from 0.100 to 0.114 $\Omega \text{ cm}^2$ and from 0.133 to 0.164 $\Omega \text{ cm}^2$ in the air containing 10% CO₂ at 750 °C during the 100 h's operation, respectively (Fig. S17a). We further investigated the electrochemical reaction evolutions of each process of the electrode impedances as a function of operation time using DRT (Fig. 3d and e). Similar to the above analyses, it is found that at least two processes in the frequency domain could be assigned to the electrode reaction (Fig. 3e), such as mass and/or gas transfer at LF, and ion and/or electron transfer at HF. According to the DRT analysis, the increase of R_p could be attributed to the sluggish surface electrode reaction of the PBSLCC and PBCC electrodes under the CO₂ poisoning condition (10% CO₂/90% air) (Fig. S17b and c). A down-shifting of a specific process on the characteristic frequency may suggest the degradation in the kinetics of the electrode process (e.g., surface reaction) due to the contaminant poisoning. The characteristic frequency of the LF peaks can be observed

to shift towards a lower frequency when the PBSLC₂₅C electrode was exposed to the air with CO₂ concentration varied from 5% to 30% (Fig. 3e). The integral resistance increased with prolonged exposure time under a higher concentration of CO₂ (e.g., 30%), strongly suggesting that high-concentration CO₂ indeed blocked the active sites for ORR and then hindered the process of the gas and mass transfer on the electrode surface (as determined by continuously increasing resistances at LF). However, when the operating conditions were in air containing 5% and 10% CO₂, the resistance of the gas and mass transfer process (LF) maintained relatively stable simultaneously with the fixed characteristic frequency, an indication of the durable high-entropy and A-site Ca-rich structure of the PBSLC₂₅C cathode against contaminant poisoning up to 10% CO₂.

3.4. Single cells performance and understanding of the CO₂ tolerance

To evaluate the ORR activity and CO₂-tolerance of the PBSLC₂₅C cathode under realistic operating conditions, a single cell with a configuration of Ni-YSZ anode support layer, Ni-YSZ anode functional layer, YSZ electrolyte (~7 μm), GDC barrier layer (~3 μm), and porous PBSLC₂₅C cathode was constructed (Fig. S18). The GDC barrier layer was utilized to prevent adverse phase reactions between the YSZ electrolyte and PBSLC₂₅C electrode, as checked by the XRD patterns (Fig. S19). Not surprisingly, the performance of the cell with the PBSLC₂₅C cathode was indeed impacted by the concentrations of CO₂.

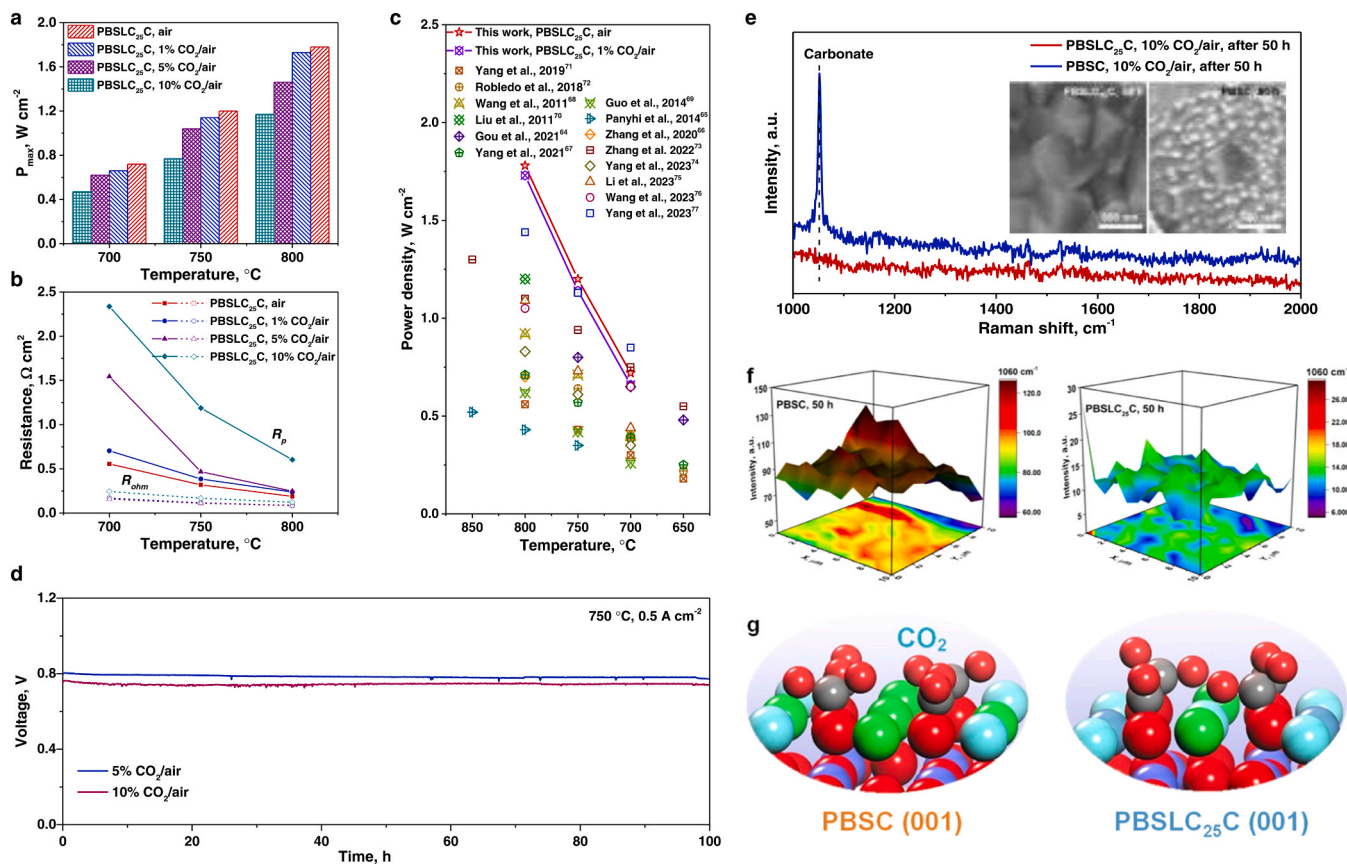


Fig. 4. Electrochemical performance of single cells with PBSLC₂₅C cathode in the air with CO₂, and understandings of high-entropy stabilized surface structure. (a) Comparisons on the peak power densities (P_{max}) and (b) resistances (R_p and R_{ohm}) of the cells with the PBSLC₂₅C cathode in air containing different concentrations of CO₂ at 700–800 °C; (c) Performance comparisons of the cell with the PBSLC₂₅C cathode with different cathode materials, [64–77] PBSLC₂₅C operating at 700–800 °C in air with/without 1% CO₂; (d) Short-term operational stabilities of the cells with the PBSLC₂₅C cathode at 750 °C under a current density of 0.5 A cm⁻² in air containing 5% and 10% CO₂; (e) Raman spectra of PBSLC₂₅C and PBSC dense membranes after exposure to air with 10% CO₂ for 50 h; insets are SEM (surface) of the as-treated PBSLC₂₅C (left) and PBSC (right) membranes in air containing 10% CO₂ for 50 h; (f) Raman mapping of SrCO₃ (peaks at 1060 cm⁻¹) on the PBSLC₂₅C (right) and PBSC (left) dense membranes after exposure to air containing 10% CO₂ for 50 h at 750 °C; and (g) Illustration of representative CO₂ adsorption on A-site cation terminated PBSC(001) (top) and PBSLC₂₅C(001) (bottom) with $\theta = 1.0$ ML (monolayer).

For instance, the P_{\max} of 1.20, 1.14, 1.04, and 0.77 W cm⁻² can be delivered at 750 °C when exposed to the ambient air and the air with 1%, 5%, and 10% CO₂, respectively (Fig. 4a, and Figs. S20–S23). Based on the comparisons of the corresponding resistances (R_{ohm} , and R_p) in Fig. 4b, the main contribution to the cell resistance increase is the R_p change of the cathode in the air with different CO₂ concentrations. The ohmic resistances (R_{ohm}) of the single cells had shown an insignificant difference. For instance, the R_{ohm} of the single cells were 0.12, 0.12, 0.12, and 0.18 Ω cm² at 750 °C in air with 0%, 1%, 5%, and 10% CO₂ respectively. The slight increase in R_{ohm} may be caused by the formation of insulating phases in the cathode/electrolyte interface or the degradation of the current collector layer.[78,79] The R_p values of the cell with the PBSLC₂₅C cathode dramatically increased from 0.39 to 1.19 Ω cm² with the CO₂ concentrations of 1–10% at 750 °C, suggesting the necessity of the improvement on the CO₂-tolerance for the developed cathodes towards actual commercialization. It is comparatively found that the P_{\max} of the cell with the PBSLC₂₅C cathode can surpass most of the previously reported high-performance SOFCs, even in the air containing CO₂ (1%) contaminant (as summarized in Fig. 4c and Table S6) [64–77]. To demonstrate the excellent CO₂-tolerance of the high-entropy PBSLC₂₅C cathode (Supplementary Note 5), we further conducted the short-term stabilities of the cell with the PBSLC₂₅C cathode at a current density of 0.5 A cm⁻² at 750 °C using the humidified H₂ (~3% H₂O) as a fuel and the air containing 5% and 10% CO₂ as an oxidant (Fig. 4d). As clearly shown, the cell with the PBSLC₂₅C cathode delivered a stable output under the conditions containing high-concentration CO₂, exhibiting excellent poisoning tolerance against 5% and 10% CO₂, which are consistent with the above symmetrical cell stabilities in Fig. 3c. The short-term stability and EIS of the single cell with the PBSLC₂₅C electrode were measured under different current densities of 0.25, 0.5, 0.75, and 1.0 A cm⁻² in air containing 5% CO₂ at 750 °C (Fig. S24). The cell with the PBSLC₂₅C electrode can maintain relatively stable voltage curves under the different applied current densities (Fig. S24a). The total resistance of the cell with PBSLC₂₅C electrode were 0.54, 0.31, 0.30, 0.29, and 0.28 Ω cm² at the current densities of 0 (OCV condition), 0.25, 0.5, 0.75, and 1.0 A cm⁻², respectively (Fig. S24b). According to the DRT analysis (Fig. S24c), peak 4 (P4) at low frequency was more significantly affected by the applied current densities when the electrode was exposed to the air containing 5% CO₂, which may correspond to the gas-associated (CO₂ and/or O₂) diffusion processes on the high-entropy electrode surface.

To gain a deeper understanding of the durable surface properties of the high-entropy PBSLC₂₅C, we have fabricated dense PBSLC₂₅C and PBSC membranes to investigate the surface interactions within the 50 h's exposure in air containing 10% CO₂ using Raman spectroscopy. Shown in Fig. 4e are the Raman spectra obtained from the PBSLC₂₅C and PBSC membranes after being exposed to air containing 10% CO₂ at 750 °C for 50 h. The as-treated PBSC showed an apparent peak near 1060 cm⁻¹, highly associated with carbonates [11] (e.g., SrCO₃ or/and BaCO₃) (Fig. S25, Supplementary Note 6). However, the Raman spectra of the PBSLC₂₅C membrane were featureless after the same treatment, verifying the high-entropy structure can inhibit the surface element (Sr or/and Ba) segregation and be finally inert for the formation of carbonates even when the CO₂ concentration reaches up to 10%. The surface of the dense PBSC membrane (inset right) was covered with many nanoparticles after exposure to air/10% CO₂, while the dense PBSLC₂₅C membrane (inset left) kept a smooth surface (Fig. 4e, inset). To further investigate the information of the carbonate formed on the membrane surface, Raman mapping of the carbonate peak at 1060 cm⁻¹ was conducted within a grid of 10 μm × 10 μm, as demonstrated in Fig. 4f. The high intensity (red region) of carbonate (1060 cm⁻¹) was considerably observed on the surface of PBSC after 50 h exposure to air/10% CO₂, and the PBSC membrane became much rougher, as obtained from the three-dimensional pattern in Fig. 4f (left). However, the PBSLC₂₅C membrane exhibited the low intensity (blue region) of carbonate (1060 cm⁻¹) with a relatively flat surface (Fig. 4f, right), indicating that

the high-entropy PBSLC₂₅C membrane is much more robust against CO₂ compared to the PBSC membrane via the effective restraining of element segregation toward the electrode surface. Furthermore, an evolution on the surface structure of PBSLC₂₅C and PBSC before and after exposure to air containing 10% CO₂ for 50 h was examined by Sr 3d X-ray photoelectron spectroscopy (XPS) curves (Fig. S26 and Supplementary Note 7). According to the Ba 3d_{5/2} and Ca 2p XPS curves, the surface Ba contents of the PBSC were higher than those of the PBSLC₂₅C under the retreatments of 10% CO₂/air at 750 °C for 50 h, and the Ca contents in the high-entropy PBSLC₂₅C structure could maintain well during the same condition (Fig. S27). The more stable surface of high-entropy PBSLC₂₅C may lead to lower ohmic and polarization resistances, and enhance cell stability. The treated PBSLC₂₅C exhibited more oxygen vacancy concentrations for ORR activity compared to the PBSC sample (Fig. S28 and Supplementary Note 8). The high-entropy PBSLC₂₅C possessed the lower started desorption temperatures of CO₂ and less CO₂ desorption amount compared to PBSC, as verified by CO₂-TPD for the powders after CO₂ pre-treatments (Fig. S29, Supplementary Note 9). These results strongly indicated that high-entropy PBSLC₂₅C with A-site Ca enrichment has an excellent structural tolerance on 10% CO₂, performing as a weaker bonding capability with CO₂ and robust surface against the carbonate formation.

Then, similar to previous studies [80–82], the outstanding CO₂ tolerance of high-entropy PBSLC₂₅C was theoretically investigated by calculating CO₂ adsorption energies in the (001) plane and compared with Sr-doped PBSC(001) (Fig. S30). We initially verified that the CO₂ adsorption on the catalytically active CoO-terminated surface is weaker bound than on A-cation-containing surfaces (Table S7). Then, we considered only A-cation-containing surfaces, similar to our previous study [80], by assuming that the peak at 1060 cm⁻¹ shown in Fig. 4e is associated with the carbonate formation resulting from segregated A-site cations from PBSC (i.e., SrO + CO₂ → SrCO₃). As schematically displayed in Figs. 4g and S30, the adsorption energy calculations support the experimental findings that the CO₂ adsorption on PBSLC₂₅C is weaker than PBSC (−0.55 eV versus −0.76 eV), which hinders the carbonate formation and maintains the excellent fuel cell performance. We theoretically verified that the non-equimolar high-entropy perovskite PBSLC₂₅C exhibits outperformed ORR activity using the bulk O p-band center and CO₂ tolerance by examining CO₂ adsorption energies.

4. Conclusions

In summary, a Ca-rich high-entropy perovskite PBSLC₂₅C cathode with excellent electrocatalytic ORR activity and CO₂ tolerance was successfully engineered and investigated under the actual cell operating conditions in both symmetrical and full cells. The PBSLC₂₅C cathode exhibited a highly active electrocatalytic performance and a remarkably low R_p of 0.032 Ω cm² at 750 °C. The analysis of the R_p durability combined with DRT suggests that the high-entropy PBSLC₂₅C can enhance the surface robustness against CO₂ poisoning, leading to highly optimized gas diffusion and surface exchange on the electrode surfaces. When applied in a full cell, the cells with the PBSLC₂₅C cathode performed exceptional peak power densities of 1.2, 1.14, 1.04, and 0.77 W cm⁻² in ambient air, and air with 1%, 5%, and 10% CO₂, respectively, at 750 °C. Excellent CO₂-tolerance with no noticeable degradations was demonstrated in air containing 5% and 10% CO₂. The high-entropy PBSLC₂₅C cathode could diminish carbonate formation and thus stabilize the perovskite structure for maintaining excellent ORR activities, as disclosed by the Raman spectroscopy and DFT calculations. Our novel high-performance electrode design strategy of high-entropy perovskites via the A-site Ca-rich can be effectively applied not only for highly CO₂-tolerance cathodes in SOFCs but also for CO₂ electrolysis, membrane reactors, oxygen separation membranes, and gas sensors.

CRediT authorship contribution statement

YongMan Choi: Writing – review & editing, Software, Resources. **Kotaro Sasaki:** Writing – review & editing. **Guangming Yang:** Data curation. **Fan He:** Writing – original draft, Visualization, Investigation, Data curation. **Yu Chen:** Writing – review & editing, Project administration, Funding acquisition, Conceptualization. **Dongliang Liu:** Investigation. **Yangsen Xu:** Investigation. **Kang Xu:** Investigation. **Feng Zhu:** Investigation.

Declaration of Competing Interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

Data availability

Data will be made available on request.

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Appendix A. Supporting information

Supplementary data associated with this article can be found in the online version at [doi:10.1016/j.apcatb.2024.124175](https://doi.org/10.1016/j.apcatb.2024.124175).

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